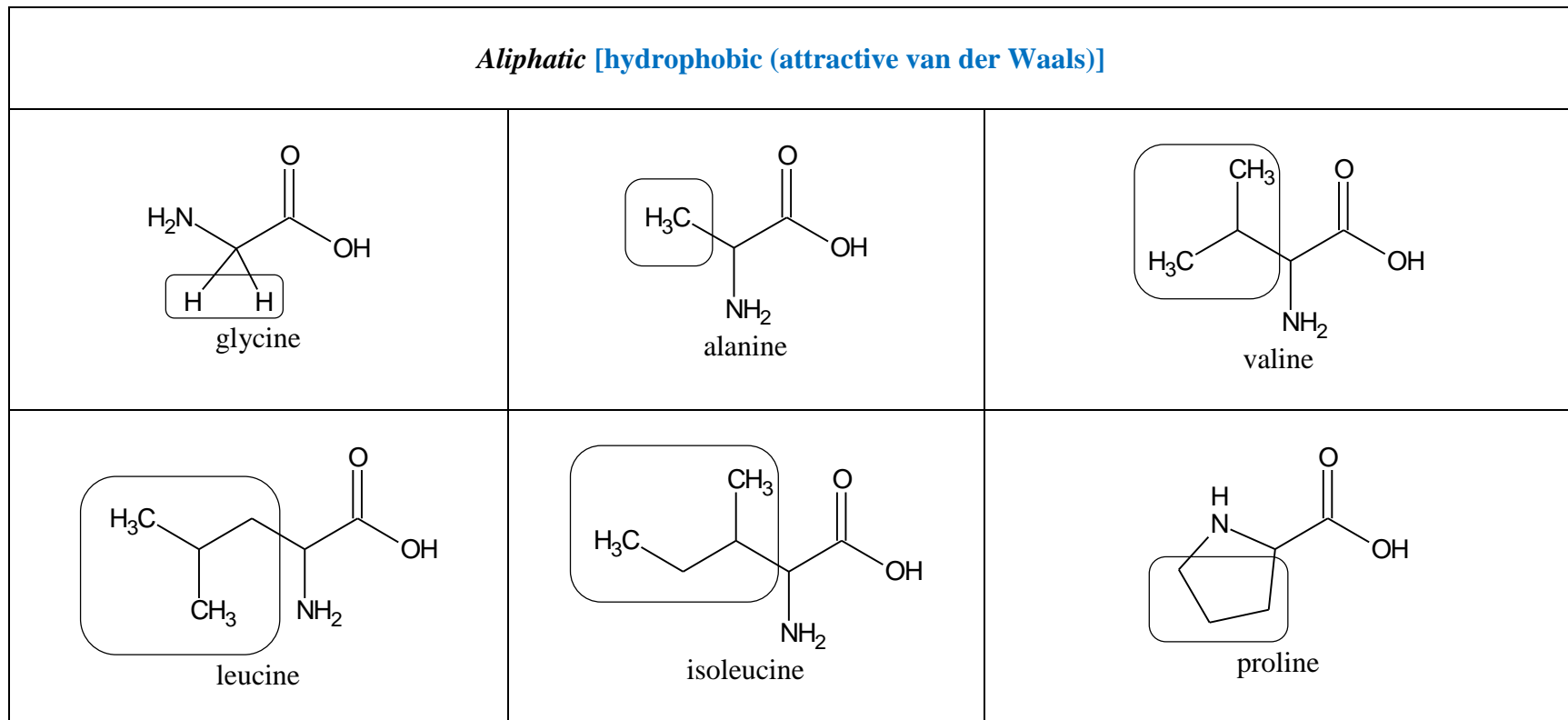


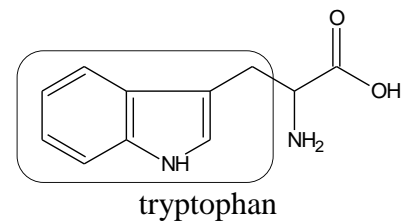
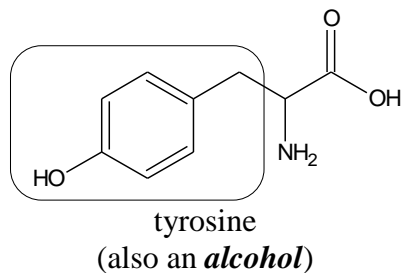
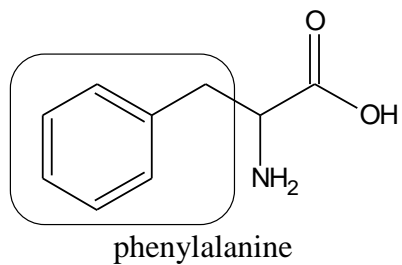
Interactions of amino acid side-groups on a protein backbone and Backbone hydrogen-bonding

The groups are organized according to their constituent atoms and the type of interaction they are likely to experience. The rectangle encloses the interacting atoms of the side groups.

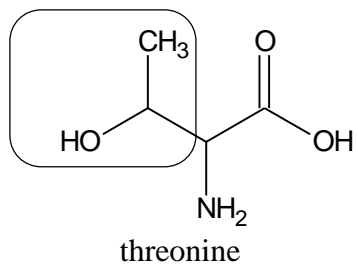
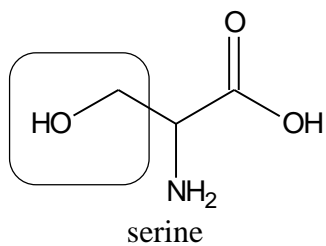
The aliphatic and aromatic groups experience weakly polar and hydrophobic interactions. Additionally, the aromatic groups can have pi-electron attractions. The alcohols and cysteine engage in dipolar/hydrogen-bonding interactions (cysteine weakly H-bonding). Methionine is hydrophobic but a potential hydrogen-bond acceptor. The amides are polar and engage in hydrogen-bonding. The charged carboxylates and aminium ions have electrostatic attractions and hydrogen-bonding.



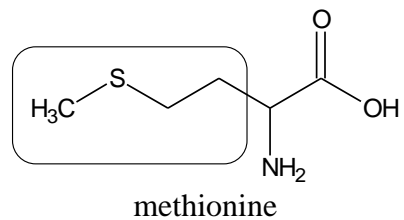
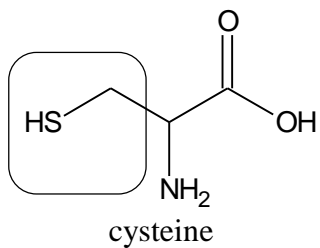
Aromatic [attractive π -interactions; hydrophobic]



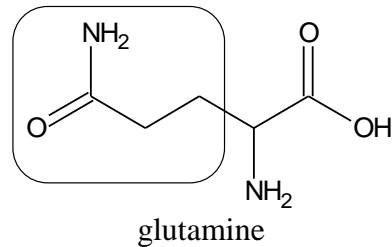
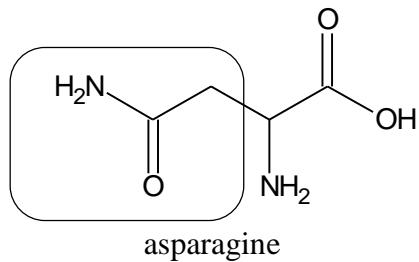
Alcohols [hydrogen-bonding; H is hydrogen-bond donor/O is hydrogen-bond acceptor]



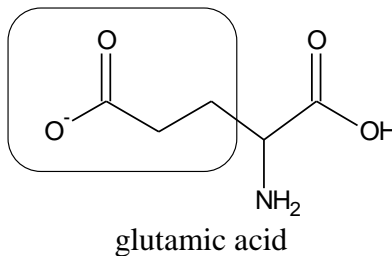
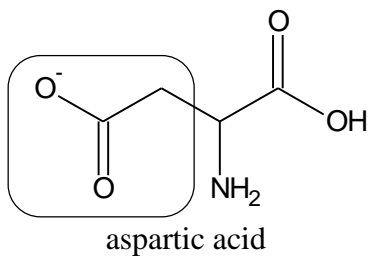
Sulfur-containing [hydrogen-bonding; H is hydrogen-bond donor/S is hydrogen-bond acceptor]



Amides [hydrogen-bonding; H is hydrogen-bond donor/O is hydrogen-bond acceptor]

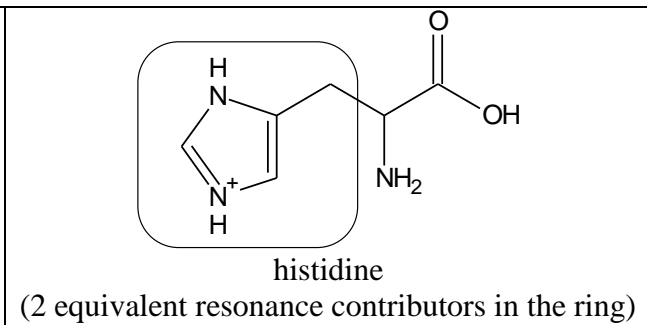
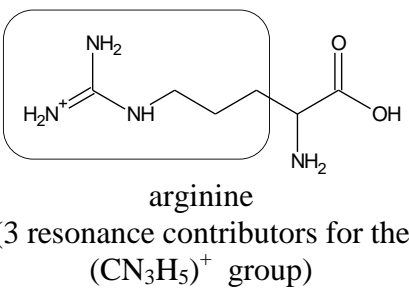
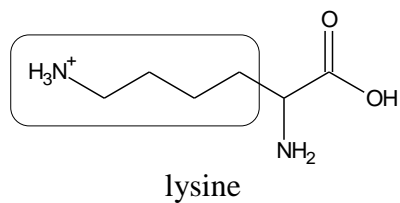


Carboxylic acids/carboxylate [hydrogen-bonding, O is hydrogen-bond acceptor; electrostatic attraction to + charge]



equivalent resonance contributors for the carboxylate (CO_2^-)

Bases/aminium ion [hydrogen-bonding, H is hydrogen-bond donor; electrostatic attraction to - charge]



The backbone itself can also engage in hydrogen-bonding interactions

